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Alcohols are comparatively more soluble in water than hydrocarbons of comparable molecular masses

Alcohols: lower alcohols are colorless liquids at normal temperature. The higher alcohols are colorless and odorless waxy solids. Phenols: phenols such as alcohols, are colorless liquids or solids. But usually they turn brown reddish due to atmospheric oxidation. Boiling Points alcohols: The alcohols of the boiling point is higher than that of the corresponding aliphatic hydrocarbons, aleolcani, aldehydes and ketones because © alcohols forming the intermolecular hydrogen bonding. The boiling point of the isomeric alcohols decreases with branching due to the decrease of the surface. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$ Δ $\text{C}_4\text{H}_9\text{OH}$ Δ $(\text{CH}_3)_2\text{CHCH}_2\text{OH}$ Δ $(\text{CH}_3)_3\text{COH}$ The boiling point of the alcohols increases with the molecular mass. Phenols: The boiling point of phenols is higher than that of the corresponding aromatic hydrocarbons, haloarenes, aldehydes and ketones because © phenols, such as alcohols, form intermolecular hydrogen bonds. Arrange 'Butan-1-ol, isobutyl alcohol, tert-butyl alcohol' Alcohol in increasing order of boiling points. © Since the boiling point decreases with branching, the boiling point increases in the order: tert-butyl alcohol

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